



**General Certificate of Secondary Education
2019–2020**

**Double Award Science
Chemistry**

Unit C1

Higher Tier

[GDW22]

THURSDAY 7 NOVEMBER 2019, MORNING

**MARK
SCHEME**

General Marking Instructions

Introduction

Mark schemes are intended to ensure that the GCSE examinations are marked consistently and fairly. The mark schemes provide markers with an indication of the nature and range of candidates' responses likely to be worthy of credit. They also set out the criteria which they should apply in allocating marks to candidates' responses.

Assessment objectives

Below are the assessment objectives for GCSE Double Award Science.

Candidates must:

- AO1** Demonstrate knowledge and understanding of:
- scientific ideas; and
 - scientific techniques and procedures;
- AO2** Apply knowledge and understanding of and develop skills in:
- scientific ideas; and
 - scientific enquiry, techniques and procedures; and
- AO3** Analyse scientific information and ideas to:
- interpret and evaluate;
 - make judgements and draw conclusions; and
 - develop and improve experimental procedures.

Quality of candidates' responses

In marking the examination papers, examiners should be looking for a quality of response reflecting the level of maturity which may reasonably be expected of a 16-year-old which is the age at which the majority of candidates sit their GCSE examinations.

Flexibility in marking

Mark schemes are not intended to be totally prescriptive. No mark scheme can cover all the responses which candidates may produce. In the event of unanticipated answers, examiners are expected to use their professional judgement to assess the validity of answers. If an answer is particularly problematic, then examiners should seek the guidance of the Supervising Examiner.

Positive marking

Examiners are encouraged to be positive in their marking, giving appropriate credit for what candidates know, understand and can do rather than penalising candidates for errors or omissions. The exception to this for GCSE Double Award Science is when examiners are marking complex calculations when the Examiners are briefed to mark by error or omission. Examiners should make use of the whole of the available mark range for any particular question and be prepared to award full marks for a response which is as good as might reasonably be expected of a 16-year-old GCSE candidate.

Awarding zero marks

Marks should only be awarded for valid responses and no marks should be awarded for an answer which is completely incorrect or inappropriate.

Marking Calculations

In marking answers involving calculations, examiners should apply the 'carry error through' rule so that candidates are not penalised more than once for a computational error. To avoid a candidate being penalised, marks can be awarded where correct conclusions or inferences are made from their incorrect calculations.

Types of mark schemes

Mark schemes for tasks or questions which require candidates to respond in extended written form are marked on the basis of levels of response which take account of the quality of written communication.

Other questions which require only short answers are marked on a point for point basis with marks awarded for each valid piece of information provided.

Levels of response

In deciding which level of response to award, examiners should look for the number of indicative content points in candidate responses to ensure that the answer has been written to coincide with the question. In deciding which mark within a particular level to award to any response, quality of communication will be assessed and examiners are expected to use their professional judgement.

The following guidance is provided to assist examiners.

- ***Threshold performance:*** Response which just merits inclusion in the level and should be awarded a mark at or near the bottom of the range.
- ***High performance:*** Response which fully satisfies the level description and should be awarded a mark at or near the top of the range.

Quality of written communication

Quality of written communication is taken into account in assessing candidates' responses to all tasks and questions that require them to respond in extended written form. These tasks and questions are marked on the basis of bands of response. The description for each band of response includes reference to the quality of written communication.

For conciseness, quality of written communication is distinguished within bands of response as follows:

- Band A: Quality of written communication is excellent.
Band B: Quality of written communication is good.
Band C: Quality of written communication is basic.
Band D: Response not worthy of credit.

In interpreting these band descriptions, examiners should refer to the more detailed guidance provided below:

Band A (Excellent): Excellent reference to scientific terminology. The candidate successfully selects and uses the most appropriate form and style of writing. Relevant material is organised with a high degree of clarity and coherence. There is widespread and accurate use of appropriate specialist vocabulary. Presentation, spelling, punctuation and grammar are of a sufficiently high standard to make meaning clear.

Band B (Good): Good reference to scientific terminology. The candidate makes a reasonable selection and use of an appropriate form and style of writing. Relevant material is organised with some clarity and coherence. There is some use of appropriate specialist vocabulary. Presentation, spelling, punctuation and grammar are sufficiently competent to make meaning clear.

Band C (Basic): Basic reference to scientific terminology. The candidate makes only a limited selection and use of an appropriate form and style of writing. The organisation of material may lack clarity and coherence. There is little use of specialist vocabulary. Presentation, spelling, punctuation and grammar may be such that intended meaning is not clear.

1 (a) (i)

Symbol	Atomic number
B	5
Al	13
Ga	31
In	49
Tl	81

Both correct [1]

[1]

(ii) number of protons in (the nucleus of) an atom

[1]

(iii) boron

[1]

(b) (i) 69

[1]

(ii) 28

[1]

(c)

Ion	Number of protons	Number of electrons	Number of neutrons	Atomic number	Mass number
Al^{3+}	13	10	14	13	27
O^{2-}	8	10	8	8	16
Br^-	35	36	46	35	81

One mark per correct row (3 × [1])

[3]

AVAILABLE
MARKS

8

- 2 (a) calcium atom shown as 2, 8, 8, 2 [1]
chlorine atom shown as 2, 8, 7 [1] [2]
- (b) concentrated hydrochloric acid [1]
blue [1]
brick red [1] [3]

(c) **Indicative points**

- Correct direction of transfer – not sharing, not covalent
- Sodium loses 1 electron/sodium becomes 2,8
- Oxygen gains 2 electrons/oxygen becomes 2,8
- Two sodium atoms required
- Correct formula Na_2O
- Sodium ion Na^+
- Oxide ion O^{2-}
- Soluble (in water)
- High melting point
- Conducts electricity when molten/dissolved in water **or** does not conduct when solid
- White solid
- brittle

Or other correct

MAX 3 IPs for properties

Band	Response	Mark
A	Candidates must use appropriate scientific terms throughout to describe the bonding of calcium with chlorine, in a logical sequence, using 7–10 of the points in the indicative content. They use good spelling, punctuation and grammar and the form and style are of a high standard.	[5]–[6]
B	Candidates use 4–6 points from the indicative content to describe the bonding of calcium with chlorine using some scientific terms and some logical sequencing. They use satisfactory spelling, punctuation and grammar and the form and style are of a satisfactory standard.	[3]–[4]
C	Candidates use 2–3 of the points from the indicative content to describe the bonding of calcium with chlorine. They use limited spelling, punctuation and grammar and make little use of scientific terms. The form and style are of a limited standard.	[1]–[2]
D	Response not worthy of credit.	[0]

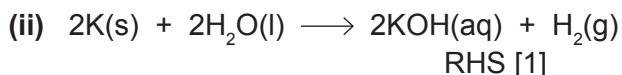
[6]

AVAILABLE
MARKS

11

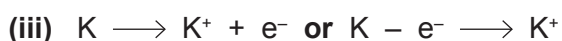
Risk	Safety precaution
burn skin/corrosive/react with water in skin [1] not harmful	Use tweezers to lift a piece of potassium
Corrosive solution splashing out	safety screen/wear (safety) goggles/ small piece of potassium [1]

[2]



correct balancing [1] – only awarded if RHS formulae both correct
correct state symbols [1] – only awarded if RHS formulae both correct

[3]



[1]

(b) explicit idea that (trend in) reactivity depends on the **outer shell of electrons** [1]

idea that outer electron is further away from the nucleus/idea of more (electron) shells [1]

idea of less tightly held outer electron/easier to lose an electron [1]

[3]

(c) they (each) lose one electron [1]

to form a positive ion/to become stable/to gain a full outer shell [1]

second mark depends on idea of electron loss

[2]

11

6 (a) potassium sulfide

[1]

(b) (i) 0.03

[1]

(ii) 0.015

Apply e.c.f. from (b)(i)

[1]

(iii) 1.65 (g) [2]

Allow 1 method mark for $M_r = 110$ [1]

Apply e.c.f. from (b)(ii)

[2]

(iv) 80% [2]

Allow 1 method mark for: $\frac{1.32}{1.65} \times 100$ or $\frac{1.32}{\text{(b)(iii) answer}} \times 100$
Apply e.c.f. from (b)(iii)

[2]

(v) reaction did not go to completion/side reactions

or other correct, e.g. loss in transfer, purity of reactant

[1]

8

				AVAILABLE MARKS	
7	(a)	(i)	B	[1]	
		(ii)	D	[1]	
		(iii)	A	[1]	
		(iv)	C [1] solid (at room temperature)/high melting point [1] conducts electricity when solid (and liquid) [1] 2nd and 3rd marks depend on 1st mark	[3]	
	(b)	(i)	positive ions/cations and no anions [1] A regular arrangement of the positive ions (apply e.c.f. for metal atoms) [1] NOT ionic bonding delocalised electrons [1] labelling [1] – 2 or more correct	[4]	
		(ii)	layers (of positive ions) can slide/move over each other [1] without disrupting the bonding [1]	[2]	12
8	(a)	(i)	colourless to dark brown	[1]	
		(ii)	none	[1]	
	(b)	(i)	Br ₂ + 2LiI → I ₂ + 2Li Br LHS [1] RHS [1] balancing – if all formulae correct [1]	[3]	
		(ii)	iodine	[1]	
	Total				
					70